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A. R.; Kreider, R. B.; Rawson, E. S.; Smith-Ryan, A. E.; Vandusseldorp, T. A.; Willoughby, D. S.; Ziegenfuss, T. N. (2021). "Common questions and misconceptions about creatine supplementation: what does the scientific evidence really show?". *Journal of the International Society of Sports Nutrition*. 18 (13): 13. doi:10.1186/s12970-021-00412-x. PMC 7871530. PMID 33557850. Franciaux M, Poortmans JR (December 2006). "Side effects of creatine supplementation in athletes". *International Journal of Sports Physiology and Performance*. 1 (4): 311–23. doi:10.1123/ijspp.1.4.311. PMID 19124889. S2CID 21330662. Bufoni J, TW, Kreider RB, Stout JR, Greenwood M, Campbell B, Spano M, et al. (August 2007). "International Society of Sports Nutrition position stand: creatine supplementation and exercise". *Journal of the International Society of Sports Nutrition*. 4. jissn: 6. doi:10.1186/1550-2783-4-6. PMC 2048496. PMID 17908288. 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For the river, see Vedita River. 1,3,5-Triazine 1,3,5-Triazine 1,3,5-Triazine Names Preferred IUPAC name 1,3,5-Triazine[1] Other names sym-Triazines-TriazineCyanidineHydrogen cyanide trimerVedita Identifiers CAS Number 290-87-9 Y 3D model (Jsmol) Interactive image ChEBI ChEBI:30259 Y ChEMBL ChEMBL15698 Y ChemSpider 8905 Y ECHA InfoCard 100.005.481 EC Number 206-028-1 PubChem CID 9262 RTECS number XY2957000 UNII 8B5F4CM81E Y CompTox Dashboard (EPA) DTXSID7052785 InChI InChI=1S/C3H3N3/c1-4-2-6-3-5-1/h1-3H YKey: JIHQDMXYFYUGFV-UHFFFAOYSA-N YnChI=1/C3H3N3/c1-4-2-6-3-5-1/h1-3H YKey: JIHQDMXYFYUGFV-UHFFFAOYAG SMILES n1cncnc1 Properties Chemical formula C3H3N3 Molar mass 81.08 g/mol Appearance White crystalline solid Melting point 81 to 83 °C (178 to 181 °F; 354 to 356 K) Structure Molecular shape planar Dipole moment zero Hazards Occupational safety and health (OHS/OSH): Main hazards Sensitive to water GHS labelling: Pictograms Signal word Danger Hazard statements H302, H314, H315, H335, H360 Precautionary statements P201, P202, P260, P261, P264, P270, P271, P280, P281, P301+P312, P301+P330+P331, P302+P352, P303+P361+P353, P304+P340, P305+P351+P338, P308+P313, P310, P312, P321, P330, P332+P313, P362, P363, P403+P233, P405, P501 Related compounds Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa). Y verify (what is YN ?) Infobox references Chemical compound 1,3,5-Triazine, also called s-triazine, is an organic chemical compound with the formula (HCN)3. It is a six-membered heterocyclic aromatic ring, one of several isomeric triazines. s-Triazine —the “symmetric” isomer—and its derivatives are useful in a variety of applications. Symmetrical 1,3,5-triazines are prepared by trimerization of certain nitriles such as cyanogen chloride or cyanamide. Benzoguanamine (with one phenyl and 2 amino substituents) is synthesised from benzonitrile and dicyandiamide.[2] In the Pinner triazine synthesis (named after Adolf Pinner)[3] the reactants are an alkyl or aryl amidine and phosgene [4][5] Insertion of an N-H moiety into a hydrazide by a copper carbenoid, followed by treatment with ammonium chloride also gives the triazine core [6] Amine-substituted triazines called Guanamines are prepared by the condensation of cyanoguanidine with the corresponding nitrile:[7] (H2N)2C=NCN + RCN → (CNH2)2(CRN)3 As a reagent in organic synthesis, s-triazine is used as the equivalent of hydrogen cyanide (HCN). Being a solid (vs a gas for HCN), triazine is sometimes easier to handle in the laboratory. One application is in the Gattermann reaction, used to attach the formyl group to aromatic substrates.[8] N- and C-substituted triazines are used industrially. The most common derivative of 1,3,5-triazine is 1,3,5-triazine-2,4,6-triamine, commonly known as melamine or cyanuramide. Another important derivative is 1,3,5-triazine-2,4,6-triol better known as cyanuric acid. Cyanuric chloride (2,4,6-trichloro-1,3,5-triazine) is the starting point for the manufacture of many herbicides such as Simazine and atrazine. Chlorinated triazines are the basis of an important family of reactive dyes, which are covalently attached to cellulosic materials.[9] Methods for attaching reactive dyes to fibres (Cell = cellulose; R = chromophore). Triazines are also found in pharmaceutical products.[10] Nomenclature of Organic Chemistry : IUPAC Recommendations and Preferred Names 2013 (Blue Book). Cambridge: The Royal Society of Chemistry. 2014. p. 147. doi:10.1039/9781849733069-PP001. ISBN 978-0-85404-182-4. Benzoguanamine J. K. Simons and M. R. Saxton Organic Syntheses Coll. Vol. 4, p.78; Vol. 33, p.13 Article Archived 2012-07-16 at the Wayback Machine A. 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Tappe, Horst; Helmling, Walter; Mischke, Peter; Rebsamen, Karl; Reiher, Uwe; Russ, Werner; Schäfer, Ludwig; Vermeiren, Petra (2000). "Reactive Dyes". *Ullmann's Encyclopedia of Industrial Chemistry*. doi:10.1002/14356007.a22.651. ISBN 3527306730. Aksenov, A. V.; Aksenova, I. V. (2009). "Use of the ring opening reactions of 1,3,5-triazines in organic synthesis (Review)". *Chemistry of Heterocyclic Compounds*. 45 (2): 130–150. doi:10.1007/s10593-009-0243-5. S2CID 94191027. Retrieved from " 3 The following pages link to 1,3,5-Triazine External tools (link count transclusion count sorted list) - See help page for transcluding these entries Showing 30 items. View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500)Creatine (links | edit) Melamine (links | edit) Photosystem II (links | edit) Simple aromatic ring (links | edit) Triazine (links | edit) 1,3,5-triazine (redirect page) (links | edit) Trimer (chemistry) (links | edit) Atrazine (links | edit) Triazine (links | edit) Cyanuric chloride (links | edit) Polyamine (links | edit) Cyanuric triazide (links | edit) S-Triazine (redirect page) (links | edit) S-triazine (redirect page) (links | edit) Pesticide degradation (links | edit) 5-Aza-7-deazapurine (links | edit) 2,4,6-Tris(trinitromethyl)-1,3,5-triazine (links | edit) Nitrogen and Non-Protein Nitrogen's effects on Agriculture (links | edit) Melamine cyanurate (links | edit) Hydrogen cyanide trimer (redirect page) (links | edit) Vedita (redirect page) (links | edit) 1,3,5-Triazine (transclusion) (links | edit) Sym-triazine (redirect page) (links | edit) Cyanide trimer (redirect page) (links | edit) Talk:1,3,5-Triazine (transclusion) (links | edit) User:CleanupListingBot/Test Reports/WikiProject Chemistry (links | edit) User:Leyo/InfoCard/0-K (links | edit) User talk:H Padleckas/Archive 2 (links | edit) Wikipedia:WikiProject Chemicals/Organization (links | edit) Wikipedia:WikiProject Chemicals/AlphabeticalList (links | edit) Wikipedia:WikiProject Chemicals/Cleanup listing (links | edit) Wikipedia:WikiProject Chemicals/Log/2009-11-28 (links | edit) Wikipedia:WikiProject Chemicals/Log/2010-10-03 (links | edit) Wikipedia:WikiProject Chemicals/Log/2011-01-20 (links | edit) Wikipedia:WikiProject Chemicals/Log/2011-08-05 (links | edit) Wikipedia:WikiProject Chemicals/Log/2011-10-19 (links | edit) Wikipedia:WikiProject Chemistry/Lists of pages/Chembox articles (links | edit) Wikipedia:Chemical infobox/Wikipedia:WikiProject Chemistry and Template:Chembox articles (links | edit) Wikipedia talk:WikiProject Chemicals/Archive 2005 (links | edit) View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500) Retrieved from " WhatLinksHere/1,3,5-Triazine" A creatina é uma molécula que faz parte do sistema fosfagénico para a produção de energia (ATP) nas células. Este sistema é usado para gerar energia de forma rápida, por exemplo, fornecendo a energia necessária aos músculos durante levantamento de peso ou corrida. As principais fontes dietéticas são alimentos ricos em proteínas, incluindo carne, peixe e aves. Uma vez sintetizada ou ingerida, a creatina é transferida do plasma através da parede intestinal para outros tecidos por transportadores específicos de creatina localizados nos músculos esqueléticos, rins, coração, fígado e cérebro. Uma vez ingerida, a creatina atravessa a parede intestinal e chega a tecidos como músculos esqueléticos, rins, coração, fígado e cérebro.Conforme envelhecemos, produzimos menos creatina. Além disso, quem nao consome ou ingere poucos alimentos de origem animal beneficia-se mais da suplementação. A forma mais estudada é a creatina monohidratada. Composta por 88% de creatina e 12% de água, ela pode ser encontrada no formato de pó solúvel em água. Seus benefícios incluem:1) Melhoria da saúde óssea. Os mecanismos são o aumento da atividade das células osteoclásticas para formação óssea e redução da reabsorção óssea, melhorando a densidade óssea e prevenindo a osteopenia e osteoporose (Candow et al., 2022).2) Fornecimento de energia para o cérebro de forma mais rápida. As células cerebrais também usam a creatina para reciclar ADP em ATP de forma mais rápida. Estudos mostram que a suplementação de creatina por 28 dias aumenta o conteúdo deste composto em 14,6% no cérebro. Estes aumentos na creatina cerebral podem traduzir-se em melhorias no comportamento e na cognição, particularmente sob condições estressantes. A creatina também ajuda a regular o humor, tratar a depressão leve, melhorar a memória e o foco, reduzir a fadiga mental (Candow et al., 2023).3) Melhoria da recuperação após lesão, concussão, AVC. Os danos após lesão cerebral traumática (LCT) geram deficiências da função mitocondrial e aumento do estresse oxidativo, além de neuroinflamação acentuada. A creatina ajuda a melhorar a bioenergética cerebral, reduzir lactato, aprimorar a função cerebrovascular e proteger o cérebro (Erdman, 2011.). 4) Melhoria na qualidade do sono. Um estudo mostrou que homens jovens que receberam suplementos de creatina antes de experimentarem 36 horas de privação de sono tiveram melhor desempenho num teste de funcionamento executivo do que indivíduos não suplementados (McMorris et al., 2006). 5) Aumenta a massa muscular pelo aumento da atividade das células satélites, e fatores de crescimento (como IGF1), proteínas quinasas, mtorc e fatores de transcrição miogénicos (Wu et al., 2022). 6) Aumenta a potência muscular. A creatina aumenta a hidratação dos músculos, melhorando seu funcionamento, sendo recomendada para ganho de força e também prevenção da perda muscular durante dietas de emagrecimento, no processo de envelhecimento e na menopausa.